

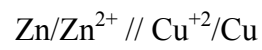
(2007/2008)

03 :

:

(12): _____

(04): _____



[Zn²⁺]=1mol/L 100mL

[Cu²⁺]=1mol/L : 100mL

Ox/Red: /1

/2

/3

$K = 1,9 \cdot 10^{37}$

. 25⁰C

: $\Delta t = 1h$

$I = 0.4A$

/4

/1 4

$\Delta t = 1h$

Cu²⁺(aq) 3 Zn²⁺(aq)

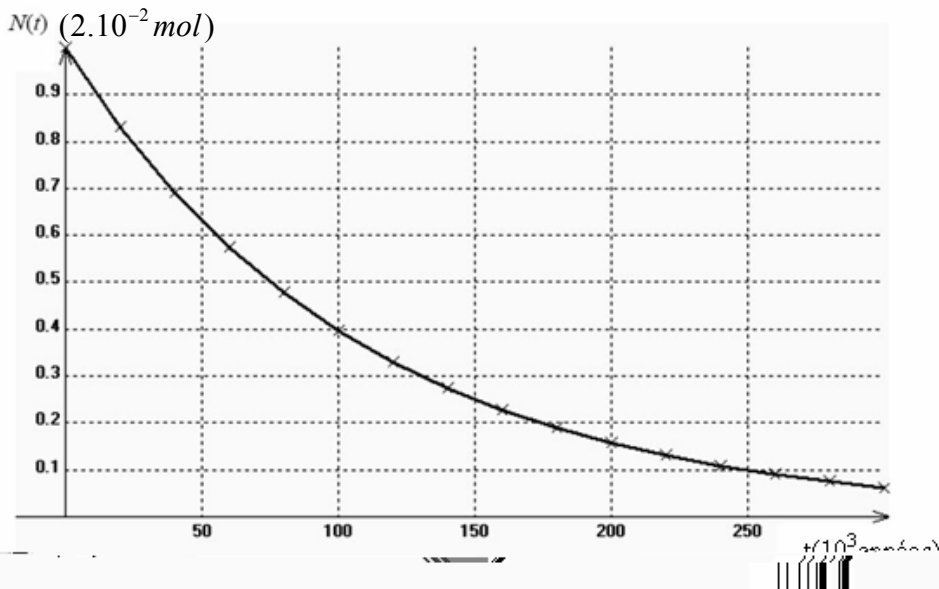
/2 4

(04): _____

.230

-1

. ²³⁰Th



${}_{88}Ra$

α

230

-2

λ

-3

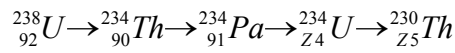
.230

-4

238

230

-5



$Z_5 \quad Z_4 :$

-

-

$\frac{N({}^{230}Th)}{N({}^{238}U)}$

230

-6

230

${}^{230}Th \quad {}^{238}U$

A(t)

:

$A(t) = -\frac{dN(t)}{dt}$

A(t)

-

$A(t) = \lambda N(t)$

$\frac{N({}^{230}Th)}{N({}^{238}U)}$

-

(04) : _____

t=0

6.5m/s

-1

$f = Kv^2$

-2

6.5m/s

-3

$g = 9.8m/s^2$

$M = 90kg$

-4

.f

$v = 2\sqrt{t}$

$0 \leq t \leq 5s$

-5

5mn

(04) : _____

() : _____

(C)

:

C

R

(1

t=0 (2)

I=660μA

(1- 2

(2-2

(3

(q)

(t)

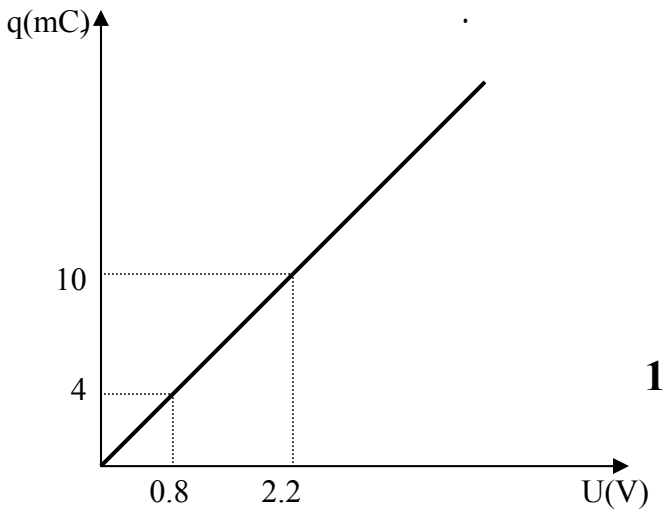
(1)

(1-3

(2-3

20% C=4700 μ F

(4



1

(04): _____
(): _____

(pictogramme)

01/01/2007 : -

31/12/2007 : -

1.02 : -

6,5 ° : -

100 g CH₃COOH

1 g 1°

03/02/2007 :

M(C) = 12 g/mol M(H) = 1 g/mol M(O) = 16 g/mol :

() :

:

$V_a = 10\text{ mL}$
 $V_b = 68\text{ mL}$:

10
 $C_b = 10^{-2}\text{ mol L}^{-1}$

:

- 1

-

-

- 2